

Diaquabis(4-methylbenzoato- κ O)bis-(nicotinamide- κ N¹)nickel(II)

Hacali Necefoğlu,^a Efdal Çimen,^a Barış Tercan,^b Emel Ermiş^c and Tuncer Hökelek^{d*}

^aDepartment of Chemistry, Kafkas University, 36100 Kars, Turkey, ^bDepartment of Physics, Karabük University, 78050 Karabük, Turkey, ^cDepartment of Chemistry, Faculty of Science, Anadolu University, 26470 Yenibağlar, Eskişehir, Turkey, and ^dDepartment of Physics, Hacettepe University, 06800 Beytepe, Ankara, Turkey

Correspondence e-mail: merzifon@hacettepe.edu.tr

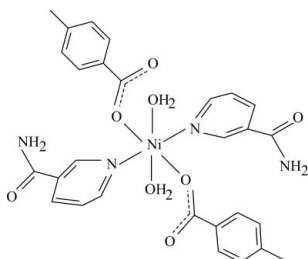
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Key indicators: single-crystal X-ray study; $T = 99$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.030; wR factor = 0.068; data-to-parameter ratio = 16.6.

The title Ni^{II} complex, $[\text{Ni}(\text{C}_8\text{H}_7\text{O}_2)_2(\text{C}_6\text{H}_6\text{N}_2\text{O})_2(\text{H}_2\text{O})_2]$, is centrosymmetric with the Ni atom located on an inversion center. The molecule contains two 4-methylbenzoate (PMB) and two nicotinamide (NA) ligands and two coordinated water molecules, all ligands being monodentate. The four O atoms in the equatorial plane around the Ni atom form a slightly distorted square-planar arrangement, while the slightly distorted octahedral coordination is completed by the two N atoms of the NA ligands in the axial positions. The dihedral angle between the carboxylate group and the adjacent benzene ring is $26.15(10)^\circ$, while the pyridine and benzene rings are oriented at a dihedral angle of $87.81(4)^\circ$. In the crystal structure, intermolecular O—H...O and N—H...O hydrogen bonds link the molecules into a three-dimensional network. The π — π contact between the benzene rings [centroid—centroid distance = $3.896(1)$ Å] may further stabilize the crystal structure. A weak C—H... π interaction involving the pyridine ring also occurs.

Related literature

For niacin, see: Krishnamachari (1974) and for the nicotinic acid derivative *N,N*-diethylnicotinamide, see: Bigoli *et al.* (1972). For related structures, see: Hökelek *et al.* (1996, 2009*a,b,c*); Hökelek & Necefoğlu (1998).



Experimental

Crystal data

$[\text{Ni}(\text{C}_8\text{H}_7\text{O}_2)_2(\text{C}_6\text{H}_6\text{N}_2\text{O})_2(\text{H}_2\text{O})_2]$	$\gamma = 71.662(2)^\circ$
$M_r = 609.26$	$V = 687.31(4)$ Å ³
Triclinic, $P\bar{1}$	$Z = 1$
$a = 7.7324(2)$ Å	Mo $K\alpha$ radiation
$b = 9.7335(3)$ Å	$\mu = 0.76$ mm ⁻¹
$c = 9.8198(3)$ Å	$T = 99$ K
$\alpha = 78.440(2)^\circ$	$0.33 \times 0.28 \times 0.25$ mm
$\beta = 86.475(3)^\circ$	

Data collection

Bruker Kappa APEXII CCD area-detector diffractometer	12002 measured reflections
Absorption correction: multi-scan (<i>SADABS</i> ; Bruker, 2005)	3390 independent reflections
$T_{\min} = 0.889$, $T_{\max} = 0.934$	3034 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.024$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.030$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.068$	$\Delta\rho_{\text{max}} = 0.48$ e Å ⁻³
$S = 1.05$	$\Delta\rho_{\text{min}} = -0.50$ e Å ⁻³
3390 reflections	
204 parameters	

Table 1

Selected bond lengths (Å).

Ni1—O1	2.0621 (10)	Ni1—N1	2.0859 (12)
Ni1—O4	2.0870 (10)		

Table 2

Hydrogen-bond geometry (Å, °).

Cg2 is the centroid of the N1/C9—C13 ring.

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N2—H21...O2 ⁱ	0.86 (2)	2.037 (19)	2.8333 (18)	153.4 (19)
N2—H22...O3 ⁱⁱ	0.90 (2)	2.05 (2)	2.9192 (19)	161.5 (18)
O4—H41...O3 ⁱⁱⁱ	0.81 (2)	2.10 (2)	2.8864 (16)	162.9 (19)
O4—H42...O2 ^{iv}	0.89 (2)	1.75 (2)	2.6240 (16)	165 (2)
C6—H6...Cg2 ^v	0.93	2.65	3.5737 (18)	171

Symmetry codes: (i) $-x, -y + 1, -z + 1$; (ii) $-x + 1, -y + 1, -z + 1$; (iii) $-x + 1, -y, -z + 1$; (iv) $-x, -y, -z + 1$; (v) $x, y, z + 1$.

Data collection: *APEX2* (Bruker, 2007); cell refinement: *SAINT* (Bruker, 2007); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 1997); software used to prepare material for publication: *WinGX* (Farrugia, 1999) and *PLATON* (Spek, 2009).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: XU2730).

References

- Bigoli, F., Braibanti, A., Pellinghelli, M. A. & Tiripicchio, A. (1972). *Acta Cryst.* **B28**, 962–966.
- Bruker (2005). *SADABS*. Bruker AXS Inc. Madison, Wisconsin, USA.
- Bruker (2007). *APEX2* and *SAINT*. Bruker AXS Inc. Madison, Wisconsin, USA.
- Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
- Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
- Hökelek, T., Dal, H., Tercan, B., Özbek, F. E. & Necefoğlu, H. (2009a). *Acta Cryst.* **E65**, m466–m467.
- Hökelek, T., Dal, H., Tercan, B., Özbek, F. E. & Necefoğlu, H. (2009b). *Acta Cryst.* **E65**, m513–m514.
- Hökelek, T., Dal, H., Tercan, B., Özbek, F. E. & Necefoğlu, H. (2009c). *Acta Cryst.* **E65**, m607–m608.
- Hökelek, T., Gündüz, H. & Necefoğlu, H. (1996). *Acta Cryst.* **C52**, 2470–2473.
- Hökelek, T. & Necefoğlu, H. (1998). *Acta Cryst.* **C54**, 1242–1244.
- Krishnamachari, K. A. V. R. (1974). *Am. J. Clin. Nutr.* **27**, 108–111.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Spek, A. L. (2009). *Acta Cryst.* **D65**, 148–155.

supplementary materials

Acta Cryst. (2010). E66, m361-m362 [doi:10.1107/S1600536810007385]

Diaquabis(4-methylbenzoato- κO)bis(nicotinamide- κN^1)nickel(II)

H. Necefoglu, E. Çimen, B. Tercan, E. Ermis and T. Hökelek

Comment

As a part of our ongoing investigation on transition metal complexes of nicotinamide (NA), one form of niacin (Krishnamachari, 1974), and/or the nicotinic acid derivative *N,N*-diethylnicotinamide (DENA), an important respiratory stimulant (Bigoli *et al.*, 1972), the title compound was synthesized and its crystal structure is reported herein.

The title complex, (I), is a crystallographically centrosymmetric mononuclear complex, consisting of two nicotinamide (NA) and two 4-methylbenzoate (PMB) ligands and two coordinated water molecules. The crystal structures of similar complexes of Cu^{II}, Co^{II}, Ni^{II}, Mn^{II} and Zn^{II} ions, [Cu(C₇H₅O₂)₂(C₁₀H₁₄N₂O)₂], (II) (Hökelek *et al.*, 1996), [Co(C₆H₆N₂O)₂(C₇H₄NO₄)₂(H₂O)₂], (III) (Hökelek & Necefoglu, 1998), [Ni(C₇H₄ClO₂)₂(C₆H₆N₂O)₂(H₂O)₂], (IV) (Hökelek *et al.*, 2009a), [Mn(C₇H₄ClO₂)₂(C₁₀H₁₄N₂O)₂(H₂O)₂], (V) (Hökelek *et al.*, 2009b) and [Zn(C₇H₄BrO₂)₂(C₆H₆N₂O)₂(H₂O)₂], (VI) (Hökelek *et al.*, 2009c) have also been reported. In (II), the two benzoate ions are coordinated to the Cu atom as bidentate ligands, while in the other structures all ligands being monodentate.

The title complex, [Ni(PMB)₂(NA)₂(H₂O)₂], has a centre of symmetry and Ni^{II} ion is surrounded by two PMB and two NA ligands and two water molecules (Fig. 1). All ligands are monodentate. The four O atoms (O1, O4, and the symmetry-related atoms, O1', O4') in the equatorial plane around the Ni atom form a slightly distorted square-planar arrangement, while the slightly distorted octahedral coordination is completed by the two N atoms of the NA ligands (N1, N1') in the axial positions (Fig. 1).

The near equality of the C1—O1 [1.2678 (17) Å] and C1—O2 [1.2654 (17) Å] bonds in the carboxylate group indicates a delocalized bonding arrangement, rather than localized single and double bonds. The average Ni—O bond length is 2.0746 (10) Å (Table 1) and the Ni atom is displaced out of the least-squares plane of the carboxylate group (O1/C1/O2) by 0.5087 (1) Å. The dihedral angle between the planar carboxylate group and the benzene ring A (C2—C7) is 26.15 (10)°, while that between rings A and B (N1/C8—C12) is 87.81 (4)°.

In the crystal structure, intermolecular O—H...O and N—H...O hydrogen bonds (Table 2) link the molecules into a three-dimensional network, in which they may be effective in the stabilization of the structure. The π — π contact between the benzene rings, Cg1—Cg1ⁱ, [symmetry code (i): 1 - x, -y, 2 - z, where Cg1 is the centroid of ring A (C2—C7)] may further stabilize the structure, with centroid-centroid distance of 3.896 (1) Å. There also exists a weak C—H... π interaction involving the pyridine ring (Table 2).

Experimental

The title compound was prepared by the reaction of NiSO₄·6(H₂O) (1.32 g, 5 mmol) in H₂O (30 ml) and NA (1.22 g, 10 mmol) in H₂O (15 ml) with sodium 4-methylbenzoate (1.36 g, 10 mmol) in H₂O (300 ml). The mixture was filtered and set aside to crystallize at ambient temperature for one week, giving blue single crystals.

Refinement

Atoms H21, H22 (for NH₂) and H41, H42 (for H₂O) were located in a difference Fourier map and refined isotropically. The remaining H atoms were positioned geometrically with C—H = 0.93 and 0.96 Å, for aromatic and methyl H atoms, respectively, and constrained to ride on their parent atoms, with $U_{\text{iso}}(\text{H}) = xU_{\text{eq}}(\text{C})$, where $x = 1.5$ for methyl H and $x = 1.2$ for aromatic H atoms.

Figures

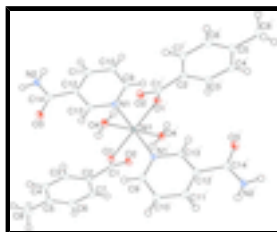


Fig. 1. The molecular structure of the title molecule with the atom-numbering scheme. Displacement ellipsoids are drawn at the 20% probability level. Primed atoms are generated by the symmetry operator: (') $-x, -y, 1 - z$.

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Crystal data

$[\text{Ni}(\text{C}_8\text{H}_7\text{O}_2)_2(\text{C}_6\text{H}_6\text{N}_2\text{O})_2(\text{H}_2\text{O})_2]$	$Z = 1$
$M_r = 609.26$	$F(000) = 318$
Triclinic, $P\bar{1}$	$D_x = 1.472 \text{ Mg m}^{-3}$
Hall symbol: $-P 1$	Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$
$a = 7.7324 (2) \text{ \AA}$	Cell parameters from 6458 reflections
$b = 9.7335 (3) \text{ \AA}$	$\theta = 2.8\text{--}28.3^\circ$
$c = 9.8198 (3) \text{ \AA}$	$\mu = 0.76 \text{ mm}^{-1}$
$\alpha = 78.440 (2)^\circ$	$T = 99 \text{ K}$
$\beta = 86.475 (3)^\circ$	Prism, blue
$\gamma = 71.662 (2)^\circ$	$0.33 \times 0.28 \times 0.25 \text{ mm}$
$V = 687.31 (4) \text{ \AA}^3$	

Data collection

Bruker Kappa APEXII CCD area-detector diffractometer	3390 independent reflections
Radiation source: fine-focus sealed tube graphite	3034 reflections with $I > 2\sigma(I)$
φ and ω scans	$R_{\text{int}} = 0.024$
Absorption correction: multi-scan (SADABS; Bruker, 2005)	$\theta_{\text{max}} = 28.4^\circ$, $\theta_{\text{min}} = 2.1^\circ$
$T_{\text{min}} = 0.889$, $T_{\text{max}} = 0.934$	$h = -10 \rightarrow 10$
12002 measured reflections	$k = -12 \rightarrow 12$
	$l = -12 \rightarrow 13$

Refinement

Refinement on F^2	Primary atom site location: structure-invariant direct methods
Least-squares matrix: full	Secondary atom site location: difference Fourier map
$R[F^2 > 2\sigma(F^2)] = 0.030$	Hydrogen site location: inferred from neighbouring sites
$wR(F^2) = 0.068$	H atoms treated by a mixture of independent and constrained refinement
$S = 1.05$	$w = 1/[\sigma^2(F_o^2) + (0.0238P)^2 + 0.4565P]$
3390 reflections	where $P = (F_o^2 + 2F_c^2)/3$
204 parameters	$(\Delta/\sigma)_{\max} < 0.001$
0 restraints	$\Delta\rho_{\max} = 0.48 \text{ e } \text{\AA}^{-3}$
	$\Delta\rho_{\min} = -0.50 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
Ni1	0.0000	0.0000	0.5000	0.00996 (8)
O1	0.12751 (13)	0.01072 (11)	0.67419 (11)	0.0129 (2)
O2	-0.11125 (14)	0.15789 (11)	0.77116 (11)	0.0149 (2)
O3	0.44171 (14)	0.33247 (11)	0.48634 (12)	0.0177 (2)
O4	0.25984 (15)	-0.07089 (12)	0.41538 (12)	0.0136 (2)
H41	0.342 (3)	-0.137 (2)	0.458 (2)	0.034 (6)*
H42	0.228 (3)	-0.103 (2)	0.345 (2)	0.036 (6)*
N1	-0.00159 (16)	0.21391 (13)	0.41144 (13)	0.0116 (2)
N2	0.33786 (19)	0.55995 (15)	0.35530 (15)	0.0176 (3)
H21	0.252 (3)	0.628 (2)	0.307 (2)	0.022 (5)*
H22	0.425 (3)	0.586 (2)	0.390 (2)	0.024 (5)*
C1	0.05766 (19)	0.09453 (16)	0.75988 (15)	0.0120 (3)
C2	0.18258 (19)	0.12691 (15)	0.85126 (15)	0.0121 (3)
C3	0.3601 (2)	0.12032 (16)	0.81040 (16)	0.0139 (3)
H3	0.4070	0.0851	0.7302	0.017*
C4	0.4678 (2)	0.16598 (16)	0.88863 (16)	0.0153 (3)
H4	0.5855	0.1624	0.8593	0.018*

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C5	0.4019 (2)	0.21690 (17)	1.00998 (17)	0.0162 (3)
C6	0.2262 (2)	0.21841 (18)	1.05279 (17)	0.0192 (3)
H6	0.1816	0.2488	1.1356	0.023*
C7	0.1171 (2)	0.17547 (17)	0.97424 (16)	0.0164 (3)
H7	-0.0006	0.1790	1.0037	0.020*
C8	0.5187 (2)	0.26752 (19)	1.09414 (19)	0.0239 (4)
H8A	0.4556	0.3664	1.1063	0.036*
H8B	0.6314	0.2651	1.0463	0.036*
H8C	0.5435	0.2033	1.1834	0.036*
C9	-0.13817 (19)	0.31071 (16)	0.33131 (16)	0.0140 (3)
H9	-0.2385	0.2823	0.3163	0.017*
C10	-0.1356 (2)	0.45107 (17)	0.27015 (17)	0.0169 (3)
H10	-0.2326	0.5157	0.2153	0.020*
C11	0.0139 (2)	0.49377 (16)	0.29208 (16)	0.0156 (3)
H11	0.0188	0.5875	0.2521	0.019*
C12	0.15612 (19)	0.39452 (16)	0.37462 (15)	0.0121 (3)
C13	0.14248 (19)	0.25568 (16)	0.43189 (15)	0.0115 (3)
H13	0.2378	0.1888	0.4869	0.014*
C14	0.32352 (19)	0.42736 (16)	0.40920 (16)	0.0131 (3)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Ni1	0.00972 (13)	0.00810 (13)	0.01255 (14)	-0.00273 (10)	-0.00152 (9)	-0.00267 (10)
O1	0.0132 (5)	0.0108 (5)	0.0144 (5)	-0.0023 (4)	-0.0020 (4)	-0.0036 (4)
O2	0.0120 (5)	0.0152 (5)	0.0175 (6)	-0.0027 (4)	-0.0015 (4)	-0.0050 (4)
O3	0.0162 (5)	0.0107 (5)	0.0257 (6)	-0.0035 (4)	-0.0068 (4)	-0.0016 (5)
O4	0.0114 (5)	0.0124 (5)	0.0163 (6)	-0.0018 (4)	-0.0020 (4)	-0.0038 (5)
N1	0.0122 (6)	0.0100 (6)	0.0128 (6)	-0.0029 (5)	-0.0008 (5)	-0.0035 (5)
N2	0.0172 (7)	0.0113 (6)	0.0254 (8)	-0.0066 (5)	-0.0074 (6)	-0.0003 (6)
C1	0.0134 (7)	0.0098 (7)	0.0124 (7)	-0.0044 (5)	-0.0023 (5)	0.0010 (5)
C2	0.0132 (7)	0.0087 (6)	0.0133 (7)	-0.0021 (5)	-0.0032 (5)	-0.0007 (5)
C3	0.0157 (7)	0.0130 (7)	0.0124 (7)	-0.0033 (6)	-0.0010 (5)	-0.0028 (6)
C4	0.0135 (7)	0.0139 (7)	0.0179 (8)	-0.0040 (6)	-0.0021 (6)	-0.0014 (6)
C5	0.0180 (7)	0.0116 (7)	0.0182 (8)	-0.0024 (6)	-0.0062 (6)	-0.0027 (6)
C6	0.0193 (8)	0.0224 (8)	0.0162 (8)	-0.0028 (6)	-0.0002 (6)	-0.0101 (7)
C7	0.0132 (7)	0.0185 (8)	0.0167 (8)	-0.0030 (6)	-0.0003 (6)	-0.0046 (6)
C8	0.0219 (8)	0.0243 (9)	0.0285 (10)	-0.0062 (7)	-0.0070 (7)	-0.0111 (7)
C9	0.0116 (7)	0.0140 (7)	0.0173 (8)	-0.0039 (6)	-0.0022 (6)	-0.0045 (6)
C10	0.0146 (7)	0.0126 (7)	0.0207 (8)	-0.0008 (6)	-0.0058 (6)	-0.0003 (6)
C11	0.0173 (7)	0.0092 (7)	0.0199 (8)	-0.0040 (6)	-0.0024 (6)	-0.0013 (6)
C12	0.0129 (7)	0.0102 (7)	0.0138 (7)	-0.0028 (5)	-0.0005 (5)	-0.0050 (6)
C13	0.0122 (6)	0.0103 (7)	0.0120 (7)	-0.0025 (5)	-0.0019 (5)	-0.0030 (5)
C14	0.0141 (7)	0.0118 (7)	0.0142 (7)	-0.0032 (6)	0.0000 (5)	-0.0053 (6)

Geometric parameters (\AA , $^\circ$)

Ni1—O1	2.0621 (10)	C4—C5	1.389 (2)
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Ni1—O1 ⁱ	2.0621 (10)	C4—H4	0.9300
Ni1—O4 ⁱ	2.0870 (10)	C5—C6	1.394 (2)
Ni1—O4	2.0870 (10)	C5—C8	1.509 (2)
Ni1—N1	2.0859 (12)	C6—H6	0.9300
Ni1—N1 ⁱ	2.0859 (12)	C7—C2	1.393 (2)
O1—C1	1.2678 (17)	C7—C6	1.383 (2)
O2—C1	1.2654 (17)	C7—H7	0.9300
O3—C14	1.2392 (18)	C8—H8A	0.9600
O4—H41	0.81 (2)	C8—H8B	0.9600
O4—H42	0.88 (2)	C8—H8C	0.9600
N1—C9	1.3421 (19)	C9—C10	1.383 (2)
N1—C13	1.3386 (18)	C9—H9	0.9300
N2—C14	1.3294 (19)	C10—H10	0.9300
N2—H21	0.86 (2)	C11—C10	1.388 (2)
N2—H22	0.89 (2)	C11—H11	0.9300
C1—C2	1.500 (2)	C12—C11	1.388 (2)
C2—C3	1.392 (2)	C12—C13	1.390 (2)
C3—H3	0.9300	C13—H13	0.9300
C4—C3	1.389 (2)	C14—C12	1.500 (2)
O1 ⁱ —Ni1—O1	180.0	C5—C4—C3	120.87 (14)
O1—Ni1—O4	86.71 (4)	C5—C4—H4	119.6
O1 ⁱ —Ni1—O4	93.29 (4)	C4—C5—C6	118.33 (14)
O1—Ni1—O4 ⁱ	93.29 (4)	C4—C5—C8	120.73 (14)
O1 ⁱ —Ni1—O4 ⁱ	86.71 (4)	C6—C5—C8	120.93 (14)
O1—Ni1—N1	89.94 (4)	C5—C6—H6	119.4
O1 ⁱ —Ni1—N1	90.06 (4)	C7—C6—C5	121.11 (14)
O1—Ni1—N1 ⁱ	90.06 (4)	C7—C6—H6	119.4
O1 ⁱ —Ni1—N1 ⁱ	89.94 (4)	C2—C7—H7	119.8
O4 ⁱ —Ni1—O4	180.0	C6—C7—C2	120.34 (14)
N1—Ni1—O4	86.75 (4)	C6—C7—H7	119.8
N1 ⁱ —Ni1—O4	93.25 (4)	C5—C8—H8A	109.5
N1—Ni1—O4 ⁱ	93.25 (4)	C5—C8—H8B	109.5
N1 ⁱ —Ni1—O4 ⁱ	86.75 (4)	C5—C8—H8C	109.5
N1 ⁱ —Ni1—N1	180.0	H8A—C8—H8B	109.5
C1—O1—Ni1	125.51 (9)	H8A—C8—H8C	109.5
Ni1—O4—H41	122.0 (15)	H8B—C8—H8C	109.5
Ni1—O4—H42	97.1 (14)	N1—C9—C10	122.60 (14)
H42—O4—H41	108 (2)	N1—C9—H9	118.7
C9—N1—Ni1	123.09 (10)	C10—C9—H9	118.7
C13—N1—Ni1	118.67 (10)	C9—C10—C11	118.88 (14)
C13—N1—C9	118.21 (12)	C9—C10—H10	120.6
C14—N2—H21	122.3 (13)	C11—C10—H10	120.6
C14—N2—H22	117.1 (12)	C10—C11—C12	119.04 (14)
H21—N2—H22	118.3 (18)	C10—C11—H11	120.5
O1—C1—C2	118.42 (13)	C12—C11—H11	120.5
O2—C1—O1	124.67 (14)	C11—C12—C13	118.29 (13)

supplementary materials

O2—C1—C2	116.86 (13)	C11—C12—C14	124.54 (13)
C3—C2—C1	121.08 (13)	C13—C12—C14	117.16 (13)
C3—C2—C7	118.83 (14)	N1—C13—C12	122.98 (13)
C7—C2—C1	119.91 (13)	N1—C13—H13	118.5
C2—C3—H3	119.8	C12—C13—H13	118.5
C4—C3—C2	120.46 (14)	O3—C14—N2	122.43 (14)
C4—C3—H3	119.8	O3—C14—C12	119.86 (13)
C3—C4—H4	119.6	N2—C14—C12	117.70 (13)
O4—Ni1—O1—C1	153.43 (11)	O2—C1—C2—C7	23.4 (2)
O4 ⁱ —Ni1—O1—C1	-26.57 (11)	C1—C2—C3—C4	172.95 (13)
N1—Ni1—O1—C1	66.68 (11)	C7—C2—C3—C4	-2.1 (2)
N1 ⁱ —Ni1—O1—C1	-113.32 (11)	C5—C4—C3—C2	1.0 (2)
O1—Ni1—N1—C9	-144.39 (11)	C3—C4—C5—C6	1.2 (2)
O1 ⁱ —Ni1—N1—C9	35.61 (11)	C3—C4—C5—C8	-179.53 (14)
O1—Ni1—N1—C13	37.62 (11)	C4—C5—C6—C7	-2.3 (2)
O1 ⁱ —Ni1—N1—C13	-142.38 (11)	C8—C5—C6—C7	178.42 (15)
O4—Ni1—N1—C9	128.90 (12)	C6—C7—C2—C1	-174.10 (14)
O4 ⁱ —Ni1—N1—C9	-51.10 (12)	C6—C7—C2—C3	1.0 (2)
O4—Ni1—N1—C13	-49.10 (11)	C2—C7—C6—C5	1.2 (2)
O4 ⁱ —Ni1—N1—C13	130.90 (11)	N1—C9—C10—C11	0.1 (2)
Ni1—O1—C1—O2	17.6 (2)	C12—C11—C10—C9	0.0 (2)
Ni1—O1—C1—C2	-159.77 (9)	C13—C12—C11—C10	0.1 (2)
Ni1—N1—C9—C10	-178.28 (11)	C14—C12—C11—C10	-178.64 (14)
C13—N1—C9—C10	-0.3 (2)	C11—C12—C13—N1	-0.3 (2)
Ni1—N1—C13—C12	178.46 (11)	C14—C12—C13—N1	178.56 (13)
C9—N1—C13—C12	0.4 (2)	O3—C14—C12—C11	178.48 (14)
O1—C1—C2—C3	26.1 (2)	O3—C14—C12—C13	-0.3 (2)
O1—C1—C2—C7	-158.94 (14)	N2—C14—C12—C11	-1.1 (2)
O2—C1—C2—C3	-151.54 (14)	N2—C14—C12—C13	-179.83 (13)

Symmetry codes: (i) $-x, -y, -z+1$.

Hydrogen-bond geometry ($\text{\AA}, ^\circ$)

Cg2 is the centroid of the N1/C9—C13 ring.

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N2—H21 \cdots O2 ⁱⁱ	0.86 (2)	2.037 (19)	2.8333 (18)	153.4 (19)
N2—H22 \cdots O3 ⁱⁱⁱ	0.90 (2)	2.05 (2)	2.9192 (19)	161.5 (18)
O4—H41 \cdots O3 ^{iv}	0.81 (2)	2.10 (2)	2.8864 (16)	162.9 (19)
O4—H42 \cdots O2 ⁱ	0.89 (2)	1.75 (2)	2.6240 (16)	165 (2)
C6—H6 \cdots Cg2 ^v	0.93	2.65	3.5737 (18)	171

Symmetry codes: (ii) $-x, -y+1, -z+1$; (iii) $-x+1, -y+1, -z+1$; (iv) $-x+1, -y, -z+1$; (i) $-x, -y, -z+1$; (v) $x, y, z+1$.

Fig. 1

